Guided Notes 16.1 – 16.2

Arrhenius, Bronsted – Lowry acid base systems, Conjugate Acid Base systems

Learning target: Explain the strengths of acids and bases using molecular structure intermolecular forces and solution equilibrium.

Define the following acids and bases and give an example

Bronsted – Lowry Acid

Bronsted – Lowry Base

Arrhenius Acid

Arrhenius Base

Conjugate Acid

Conjugate Base

What is amphoteric and give an example of an amphoteric substance.

What is an oxyacid, binary acid?

When is an oxyacid strong? Give examples

Give examples of strong binary acids

Give an example of a weak acid

Give example of a strong Base

Give an example of a weak base

What is a strong Acid, how does it dissociate, which direction does the equilibrium lie

What is a weak acid, how does it dissociate which direction does the equilibrium lie

What is a strong base, how does it dissociate which way does the equilibrium lie

What is a weak base, how does it dissociate which direction does the equilibrium lie.

Try sample exercise 16.1-16.3 and the practice exercises

See Go figure on page 653 Brown LeMay

Draw and analyze the 16.3 reactions on page 653

Using HCl show an equation representing HCl as a Bronsted Lowry model and then as a Arrhenius Acid

What two ions are central to the Arrhenius definitions of acids and bases?

In the forward of this equilibrium reaction, which substance is the Acid, Base, CA, CB

H2S + CH3NH2 ↔ HS- + CH3NH3+