

Cations		Anions	
1+	Alkali metals lithium, Li ⁺ ; sodium, Na ⁺ ; potassium, K ⁺ ; rubidium, Rb ⁺ ; cesium, Cs ⁺ ; francium, Fr ⁺ ; hydrogen, H ⁺	1-	Halogens fluoride, F ⁻ ; chloride, Cl ⁻ ; bromide, Br ⁻ ; iodide, I ⁻ ; astatide, At ⁻
2+	Alkaline Earths: beryllium, Be ²⁺ ; magnesium, Mg ²⁺ ; calcium, Ca ²⁺ ; strontium, Sr ²⁺ ; barium, Ba ²⁺ ; radium, Ra ²⁺	2-	Oxygen family oxide, O ²⁻ ; sulfide, S ²⁻ ; selenide, Se ²⁻ ; telluride, Te ²⁻ ; polonide, Po ²⁻
3+	aluminum, Al ³⁺ ; boron, B ³⁺	3-	Nitrogen family nitride, N ³⁻ ; phosphide, P ³⁻ ; arsenide, As ³⁻
4+		4-	carbide (covalent), C ⁴⁻

Summary of Cations with Variable Oxidation Numbers—Stock System

1+, 2+ copper(I), Cu⁺; copper(II), Cu²⁺;
mercury(I)*, Hg₂²⁺; mercury(II), Hg²⁺.
*Note: mercury(I) actually exists as a diatomic ion and is written as Hg₂²⁺ and not Hg⁺.

1+, 3+ gold(I), Au⁺; gold(III), Au³⁺;
indium(I), In⁺; indium(III), In³⁺;
thallium(I), Tl⁺; thallium(III), Tl³⁺;
chromium(II), Cr²⁺; chromium(III), Cr³⁺;
cobalt(II), Co²⁺; cobalt(III), Co³⁺;
iron(II), Fe²⁺; iron(III), Fe³⁺;
manganese(II), Mn²⁺; manganese(III), Mn³⁺;
lead(II), Pb²⁺; lead(IV), Pb⁴⁺;

2+, 3+ tin(II), Sn²⁺; tin(IV), Sn⁴⁺;
zirconium(II), Zr²⁺; zirconium(IV), Zr⁴⁺;
cerium(III), Ce³⁺; cerium(IV), Ce⁴⁺;
antimony(III), Sb³⁺; antimony(V), Sb⁵⁺;
arsenic(III), As³⁺; arsenic(V), As⁵⁺;
bismuth(III), Bi³⁺; bismuth(V), Bi⁵⁺;
phosphorus(III), P³⁺; phosphorus(V), P⁵⁺;
2+, 3+, 4+ iridium(I), Ir⁺; iridium(III), Ir³⁺; iridium(IV), Ir⁴⁺;
titanium(II), Ti²⁺; titanium(III), Ti³⁺; titanium(IV), Ti⁴⁺;
tungsten(II), W²⁺; tungsten(IV), W⁴⁺; tungsten(V), W⁵⁺;
3+, 4+, 5+ uranium(III), U³⁺; uranium(IV), U⁴⁺; uranium(V), U⁵⁺;
2+, 3+, 4+, 5+ vanadium(II), V²⁺; vanadium(III), V³⁺;

vanadium(IV), V⁴⁺; vanadium(V), V⁵⁺.
Note: When reading the name of an ion such as Pb²⁺, the ion is read in English as the "lead two ion."