STOICHIOMETRY TEST A

1. 115g of lithium reacts with oxygen gas. How many grams of lithium oxide will be produced?
2. 23.0g of Magnesium hydroxide is needed to make 540.0ml solution. What is the Molarity?
3. How many grams of Sodium Chlorate will be needed to make 375ml of a 3.50M solution?
4. When copper(II) replaces silver in 208ml of a .100M Silver nitrate solution, how many grams of silver will be produced?
5. 79.1g of Zn is added to 1050ml of a 2.00 M HCl solution. a) What is the LR? B) How many grams of zinc Chloride will be formed ? c) If 141g of zinc Chloride was actually produced what would be the %yield?

Matching

6\_\_\_\_solute a) dissolving meduim

 7\_\_\_\_ weak electrolyte b) does not conduct electricity

8\_\_\_\_\_ strong electrolyte c) moles/liters

9)\_\_\_\_\_ Limiting reactant d) slightly conducts electricity

10) \_\_\_\_nonelectrolyte e) determines the amount of product

11\_\_\_\_\_solvent f) substance being dissolved

12) \_\_\_\_\_ Molarity g) strong conductor of current

STOICHIOMETRY TEST B

1)115g of potassium reacts with oxygen gas. How many grams of lithium oxide will be produced?

2)23.0g of Calciium hydroxide is needed to make 540.0ml solution. What is the Molarity?

3) How many grams of Lithium Chlorate will be needed to make 375ml of a 3.50M solution?

4) When copper(II) replaces silver in 308ml of a .200M Silver nitrate solution, how many grams of silver will be produced?

5)79.1g of Zn is added to 1050ml of a 2.00 M HCl solution. a) What is the LR? B) How many grams of zinc Chloride will be formed ? c) If 140g of zinc Chloride was actually produced what would be the %yield?

Matching

6\_\_\_\_solvent a) dissolving meduim

 7\_\_\_\_ strong electrolyte b) does not conduct electricity

8\_\_\_\_\_ nonelectrolyte c) moles/liters

9)\_\_\_\_\_ Molarity d) slightly conducts electricity

10) \_\_\_\_weak electrolyte e) determines the amount of product

11\_\_\_\_\_solute f) substance being dissolved

12) \_\_\_\_\_ Limiting reactant g) strong conductor of current